Oh Pascal

question 71ce2 socraticquestion 9f499 socraticquestion c548d socraticcalculating the concentration of excess hcl in e socraticquestion e7848 socraticquestion 370a7 socraticquestion 750c8 socraticquestion 71b91 socraticquestion a721d socraticquestion d9b58 socratic www.bing.com www.bing.com www.bing.com www.bing.com www.bing.com www.bing.com www.bing.com www.bing.com www.bing.com www.bing.com

h oh h 2o when the acid was added to the resulting solution the h and oh react in a 1 1 ratio this tells us that the number of moles of h used will be equal to the number of oh moles in

explanation your starting point here is the ph of the solution more specifically you need to use the given ph to determine the concentration of hydroxide anions oh present in the saturated

question 1 k sp 1 1 xx10 11 question 2 s 4 9 xx10 12 m quest 1 determine the ksp for magnesium hydroxide mg oh 2 where the molar solubility of mg

the acid in excess is then titrated with n aoh aq of known concentration we can thus get back to the concentration or molar quantity of m oh 2 as it stands the question and answer are

similarly oh becomes h 2o indicating a gain of a h ion so you can say that nh 4 is the acid and oh is the base conjugates are basically the other term for every acid you have a

the sodium ions remain in solution as spectator ions if xs sodium hydroxide is added the precipitate redissolves to give the soluble plumbate ii ion a simple way of writing this is chemguideuk

here s what i got the problem wants you to use the base dissociation constant k b of ammonia nh 3 to determine the percent of ammonia molecules that ionize to produce ammonium cations

since water is in excess 67 7 g mgo are needed to produce 98 0 g mg oh 2 balanced equation mgo s h 2 o l rarr mg oh 2 s moles magnesium hydroxide start with the given mass of

ph 1 61151 oh 4 08797 10 13m hf 0 855538m h 0 024462m f 0 024462m hf h 2o h 3o f we can find the concentration of h or h 3o by three ways one is by the

explanation h 3po 4 aq ca oh 2 aq rarr ca 3 po 4 2 h 2o l

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